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*Date Submitted:* 2020-02-12

*Keywords:* municipal sludge, coking wastewater, adsorption kinetics, activated carbon

## *Abstract:*

The toxic pollutants phenol and cyanide in the bio-treated effluent of coking wastewater still need advanced treatment to meet environmental requirements. In this study, activated carbon prepared from municipal sludge and bamboo waste (SBAC) was used for simultaneous adsorption of phenol and cyanide from bio-treated effluent of coking wastewater. The results showed that the optimum removal efficiencies of volatile phenol (69.7%) and total cyanide (80.1%) were observed at a SBAC dosage of 8 g/L, a pH value of 8.0, and a contact time of 80 min. The physical and chemical properties of SBAC were analyzed using Brunauer-Emmett-Teller (BET) surface area (SBET), scanning electron microscopy (SEM), and Fourier transform infrared (FTIR) spectroscopy. SBAC had high SBET (289.58 m<sup>2</sup>/g) and rich mesoporous structure (average pore diameter of 3.688 nm), and carboxylic groups on SBAC surfaces were enhanced due to the addition of bamboo waste. In addition, a kinetic model of pseudo-first-order fitted well with the experimental data of volatile phenol, while the adsorption of total cyanide onto the SBAC was better described by a pseudo-second-order kinetic model.

*Record Type:* Published Article

*Submitted To:* LAPSE (Living Archive for Process Systems Engineering)

*Citation (overall record, always the latest version):*

LAPSE:2020.0209

*Citation (this specific file, latest version):*

LAPSE:2020.0209-1

*Citation (this specific file, this version):*

LAPSE:2020.0209-1v1

*DOI of Published Version:* <https://doi.org/10.3390/pr8010082>

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Article

# Application and Mechanism of Sludge-Based Activated Carbon for Phenol and Cyanide Removal from Bio-Treated Effluent of Coking Wastewater

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Received: 9 December 2019; Accepted: 6 January 2020; Published: 8 January 2020



**Abstract:** The toxic pollutants phenol and cyanide in the bio-treated effluent of coking wastewater still need advanced treatment to meet environmental requirements. In this study, activated carbon prepared from municipal sludge and bamboo waste (SBAC) was used for simultaneous adsorption of phenol and cyanide from bio-treated effluent of coking wastewater. The results showed that the optimum removal efficiencies of volatile phenol (69.7%) and total cyanide (80.1%) were observed at a SBAC dosage of 8 g/L, a pH value of 8.0, and a contact time of 80 min. The physical and chemical properties of SBAC were analyzed using Brunauer–Emmett–Teller (BET) surface area ( $S_{BET}$ ), scanning electron microscopy (SEM), and Fourier transform infrared (FTIR) spectroscopy. SBAC had high  $S_{BET}$  (289.58 m<sup>2</sup>/g) and rich mesoporous structure (average pore diameter of 3.688 nm), and carboxylic groups on SBAC surfaces were enhanced due to the addition of bamboo waste. In addition, a kinetic model of pseudo-first-order fitted well with the experimental data of volatile phenol, while the adsorption of total cyanide onto the SBAC was better described by a pseudo-second-order kinetic model.

**Keywords:** activated carbon; adsorption kinetics; coking wastewater; municipal sludge

## 1. Introduction

In China, approximately 31.2 million tons of coking wastewater were produced from coal gas purification, coal coking, and the by-products recovery process of coking in 2017 [1]. Coking wastewater contains a large amount of organic pollutants, such as phenol and cyanide, which are highly toxic to both human and aquatic life even at low concentrations [2]. Therefore, it is critical to treat coking wastewater before it is discharged into the environment.

To date, numerous methods have been developed and applied to coking wastewater treatment [3–5], however, biological technologies are still widely used in China because they are eco-friendly methods for degrading organic pollutants into harmless substances [5,6]. It is reported that almost 95–98% of the organics in coking wastewater can be removed by biological degradation, but the phenol and cyanide contents in the bio-treated effluent still need further treatment to meet strict environmental requirements [1]. Therefore, filtration, coagulation, adsorption, and ozonation are used in the advanced treatment of coking wastewater after biological treatments [6–10]. Adsorption with activated carbon (AC) has been considered one of the most attractive techniques due to its simplicity and high removal performance [11,12]. However, the application of commercial AC in wastewater treatment is limited by its high-cost and non-renewable nature, therefore, many efforts have recently been made to prepare AC from renewable, abundant, and low-cost biomass [13–15]. For example, Singh et al. [15] used copper-impregnated coconut shell AC (Cu-CSAC) in a binary aqueous solution of phenol and cyanide,

and the results showed that the optimal adsorption efficiencies for phenol (71.43%) and cyanide (86.8%) were observed at the temperature of 30 °C, the pH of 8, and the adsorbent dose of 40 g/L.

Sludge is an inevitable by-product of municipal wastewater treatment plants (WWTPs). About 13 million tons (dry weight) of municipal sludge (MS) were produced in China in 2016 [16]. Preparation of AC from MS by pyrolysis/activation is a promising method for the sustainable management of MS, since MS is a carbon-rich, renewable, and readily accessible resource [12,17]. Nevertheless, the adsorption capacities of sludge-based AC (SAC) to contaminants are usually low due to the relatively high ash and inorganics in the MS [18]. Therefore, a wide variety of agricultural wastes, nuts, and plant leaves have been added into MS to improve its surface area, internal pore structure, and the chemical functional groups present on its surface [19,20]. Recent research has found that bamboo waste (BW) is a promising type of additive in the pyrolysis of MS because of its high lignocellulose content and specific morphological properties [21,22]. The hemicellulose and cellulose in BW are pyrolyzed to produce tar at 220–315 °C and 315–400 °C, respectively [23], and some tars are continuously decomposed above 400 °C into pyrolytic gas, promoting the formation of micropores [21]. Moreover, the AC prepared from BW has expanded carbon nanotubes and rich aromatic structures [22]. Currently, SAC has showed potential in removing metals, phenol, or cyanide from single or binary aqueous solutions [24]. However, the utilization of SAC for the co-adsorption of phenol and cyanide from real wastewater has not been reported yet.

Consequently, in this study, the AC produced from MS and BW (SBAC) was used for the simultaneous adsorption of volatile phenol and total cyanide from bio-treated effluent of real coking wastewater. The optimal values for operational parameters such as SBAC dosage, initial pH, and contact time for efficient adsorption of phenol and cyanide in coking wastewater were investigated. Simultaneously, the adsorption mechanism of phenol and cyanide was determined by analyzing the physical and chemical properties of SBAC using scanning electron microscopy (SEM) and Fourier transform infrared spectrometer (FTIR). Additionally, the adsorption of phenol and cyanide onto SBAC was evaluated using adsorption kinetic models.

## 2. Materials and Methods

### 2.1. Precursor of SBAC

The MS used in this study was collected from a dewatering room of the Qiaobei WWTP in Nanjing, China. The BW was obtained from a furniture factory in Nanjing, China. The characteristics of MS and BW are shown in Table 1. The MS and BW were dried to constant weight at 105 °C using a drying oven (DGG-9030A, Donglu, Shanghai, China) and then crushed and sieved to a size below 80-mesh. Finally, these two powders were separately packed in sealed plastic bags and stored in a desiccator until use.

**Table 1.** Characteristics of municipal sludge (MS) and bamboo waste (BW).

Parameters	Volatile Solids (VS) (dw %)	Ash (dw %)	Water Content (ww %)	C (dw %)	H (dw %)	N (dw %)
MS	62.01	37.99	78.86	26.5	6.24	4.08
BW	98.64	1.36	9.94	48.5	3.35	0.75

dw: dry weight. ww: wet weight.

### 2.2. Preparation of SBAC

The dry MS and BW were mixed well in a stainless steel reactor with mass ratio of 2:1, and then soaked in 5 mol/L of ZnCl<sub>2</sub> solution (ZnCl<sub>2</sub>: raw material = 3:1, *w/w*) for 24 h. Next, the mixture was centrifuged at 1176× *g* for 3 min. Afterwards, the sediment was dried at 105 °C for 12 h, carbonized at 550 °C for 30 min in a resistance furnace (SX-12-10, Xinyi, Shanghai, China), and cooled to room temperature overnight. Next, the carbonized product was immersed in 1.0 mol/L HCl for 5 min at

room temperature, and the above step was repeated three times. Finally, the product was rinsed to neutral pH with deionized water, dried, and pulverized to below 80-mesh to obtain SBAC.

The preparation steps of SAC were similar to that of SBAC, except that no BW was added to the MS.

### 2.3. Characteristics of SBAC

The porosity and surface area of SAC and SBAC samples were measured by a surface area pore analyzer (Micrometrics, ASAP2020, Norcross, GA, USA). The Brunauer–Emmett–Teller (BET) surface area ( $S_{\text{BET}}$ ) was calculated by use of the BET equation. The total pore volume ( $V_{\text{T}}$ ) was determined at relative pressure of 0.995, and the micropore volume ( $V_{\text{mic}}$ ) was analyzed by the de Boer method [25]. The average pore diameter ( $D_{\text{p}}$ ) of the SBAC was calculated using the equation  $4V_{\text{T}}/S_{\text{BET}}$  [26].

The surface morphology of SAC and SBAC was measured by a SEM microscope (QUANTA 200, FEI, Hillsboro, OR, USA). The surface functional groups on the SAC and SBAC were evaluated with a FTIR spectrometer (IR-360, Nicolet, Madison, WI, USA), and the spectra were obtained in a wavelength rang of 400–4200  $\text{cm}^{-1}$ .

### 2.4. Adsorption Tests

Adsorption tests of SBAC were performed in three trials, using coking wastewater as adsorbate. The coking wastewater was taken from the bio-treated effluent in Shougang Jingtang Iron and Steel Co., Ltd. (Tangshan, China). The characteristics of the wastewater are summarized in Table 2. The coking wastewater was stored at 4 °C to avoid any changes in physical and chemical properties before use.

**Table 2.** Characteristics of coking wastewater.

Parameters	Coking Wastewater
Chemical oxygen demand (COD) (mg/L)	2457.6 ± 125.4
Ammonia nitrogen (mg/L)	316.3 ± 24.2
Total cyanide (mg/L)	35.4 ± 1.5
Volatile phenol (mg/L)	227.5 ± 7.6
Suspended solids (SS) (mg/L)	76.9 ± 6.5
pH	8.1 ± 0.8

Operational parameters are important variables affecting the adsorption process. The pH influences the surface charge and functional groups of adsorbent, as well as the ionization capacity of the adsorbate [27]. Contact time is another important parameter to determine the equilibrium time and kinetics of adsorption process [12]. Larger surface area and more binding sites were found with increase of adsorbent dosage, which directly affects the removal efficiency of adsorbate [28]. Therefore, the effects of different adsorption parameters such as initial pH (4–10), contact time (0–120 min), and adsorbent dosage (1–12 g/L) on adsorption process were investigated. All adsorption experiments were performed in 250.0-mL conical flasks, in each of which 100.0 mL of coking wastewater and different amounts of SBAC were placed. The solutions were shaken at 100 rpm using a thermostatic shaker incubator (SHZ-82, Changzhou Guohua Electric Appliance Co., Ltd., Changzhou, China). The initial pH values were adjusted by adding 1.0 M of hydrochloric acid or 1.0 M of sodium hydroxide solution. At the end of each experiment, samples were filtered through 0.45- $\mu\text{m}$  Millipore HA membranes (Millipore, Billerica, MA, USA), and the concentrations of volatile phenol and total cyanide in the filtrates were determined.

All experiments were conducted in triplicate, and the average and standard deviation values were recorded.

## 2.5. Analytical Methods

The COD, ammonia nitrogen, water content, VS, and SS were measured according to standard methods [29]. The pH was measured by a pH meter (PHS-3C, Leici, Shanghai, China). The elemental composition of the MS and BW (C, H, and N) was determined by an automatic elemental analyzer (Vario EL cube, Elementar, Hanau, Germany). The concentrations of volatile phenol and total cyanide in the coking wastewater were determined by 4-aminoantipyrine (20, pp. 1014) and pyridine-barbituric acid (20, pp. 607) spectrophotometry, respectively [29].

The removal efficiency ( $E$ , %) and the equilibrium adsorption capacity ( $q_e$ , mg/g) of SBAC on volatile phenol and total cyanide were calculated using Equations (1) and (2):

$$E = \frac{C_0 - C_e}{C_0} \times 100\% \quad (1)$$

$$q_e = \frac{(C_0 - C_e) \times V}{W} \quad (2)$$

where  $q_e$  represents the amount of pollutants adsorbed in mg per g of SBAC at equilibrium conditions;  $C_0$  and  $C_e$  are initial and equilibrium concentrations of volatile phenol or total cyanide (mg/L), respectively;  $V$  is the volume of coking wastewater (L); and  $W$  is the weight of SBAC (g).

## 3. Results and Discussion

### 3.1. Characterizations of the SBAC

#### 3.1.1. BET Surface Area and Pore Structure

Nitrogen ( $N_2$ ) adsorption–desorption isotherms of SAC and SBAC are shown in Figure 1. According to the International Union of Pure and Applied Chemistry (IUPAC) classification, the shape of isotherms and hysteresis loops were classified as Type II with H3 hysteresis, which indicated that the SAC and SBAC samples had mesoporous structures [30]. Simultaneously, it can be seen that the adsorbed volume of SBAC was almost twice that of the SAC, indicating that SBAC had a richer mesoporous structure than SAC.

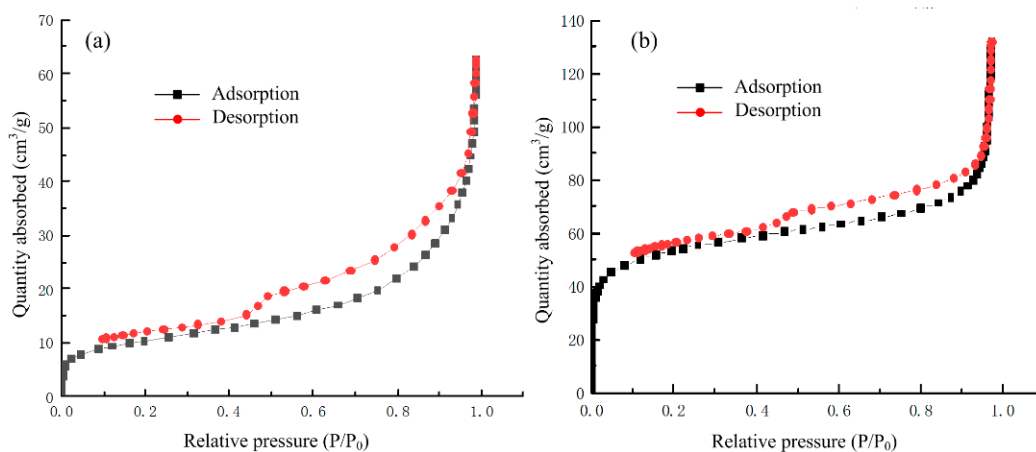


Figure 1.  $N_2$  adsorption–desorption isotherm of (a) SAC and (b) SBAC.

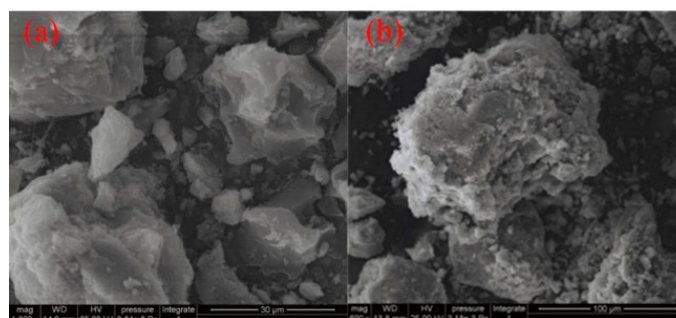
The pore structure characteristics of SAC and SBAC are listed in Table 3. The  $D_p$  of SBAC was 3.688 nm, which was much lower than that of SAC (10.941 nm). The  $S_{BET}$  of the SBAC was found to be 289.578  $m^2/g$  with a  $V_T$  of 0.267  $cm^3/g$ , which were almost 4.08 and 1.38 times higher than that of the SAC, respectively. It is argued that the addition of BW was effective in enhancing the surface area and pore volume by converting macropores or mesopores to micropores. As previously reported, when the temperature is above 400 °C, some tar produced from hemicellulose/cellulose in BW is broken down and converted into pyrolytic gas, thereby promoting the formation of pores [21]. Therefore, the addition of BW to improve carbon content in MS was the most likely reason for the significant increase in the surface area and pore volume of SBAC.

**Table 3.** Characteristics of porous structure of SAC and SBAC.

Samples	$S_{BET}$ ( $m^2/g$ )	Micropore Surface Area ( $m^2/g$ )	$V_T$ ( $cm^3/g$ )	$V_{mic}$ ( $cm^3/g$ )	$D_p$ (nm)
SAC	70.921	18.568	0.194	0.016	10.941
SBAC	289.578	70.523	0.267	0.049	3.688

### 3.1.2. SEM Analyses

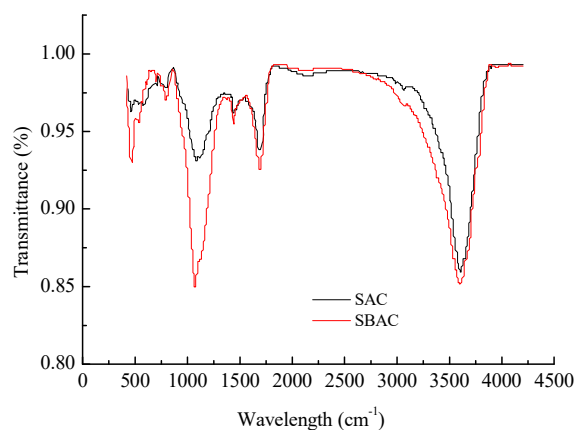
SEM analysis results of the SAC and SBAC samples are shown in Figure 2. Unlike the flaky and massive particles on the surface of the SAC, the SBAC exhibited a more irregular and looser surface morphology. These surface features might provide sufficient diffusion spaces for contaminants in the coking wastewater.



**Figure 2.** SEM images of (a) SAC and (b) SBAC.

### 3.1.3. FTIR Analysis

It is well known that the adsorption performance of AC is strongly affected by functional groups such as carboxyl, carbonyl, phenolic hydroxyl, and lactones on the surface [30,31]. FTIR spectra of the SAC and SBAC samples are shown in Figure 3. The broad absorption band in the range of 3300–3600  $cm^{-1}$  is ascribed to the O-H stretching vibrations, which correspond to associated peaks of hydroxyl groups from phenols and alcohols [19]. The distinctive peak observed at 1624  $cm^{-1}$  can be attributed to the C=O and C=C vibration of carbonyl in carboxylic or ester groups [32]. The absorption peak at 1383  $cm^{-1}$  seems to be due to the lactone groups [33]. The relatively intense band at 1042  $cm^{-1}$  can be attributable to alcohol groups (C-OH) [32]. In addition, the peaks at 465  $cm^{-1}$  represented the C-H stretching vibration [33]. In this study, the addition of BW led to an increase of oxygen-containing functional groups (C-OH, O-H, and C=O), which might contribute to the adsorption of total cyanide.



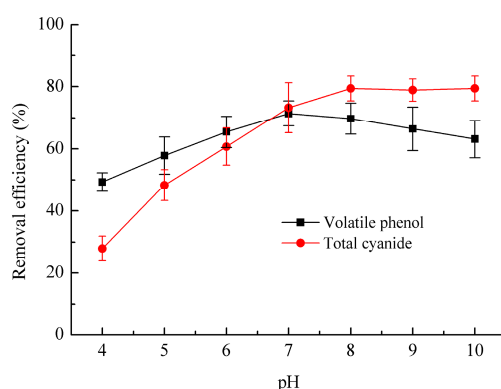
**Figure 3.** FTIR of SAC and SBAC samples.

### 3.2. Adsorption Performances of SBAC

#### 3.2.1. Effect of Initial pH

The effect of initial pH on the removal efficiencies of volatile phenol and total cyanide in coking wastewater was assessed at a SBAC dosage of 8 g/L and a contact time of 80 min. As shown in Figure 4, the removal efficiencies of volatile phenol increased at first and then decreased with the increase of pH, and the maximal removal efficiency (71.3%) was observed at pH 7.0. On the other hand, as the pH increased from 4.0 to 8.0, the removal efficiencies of total cyanide increased rapidly from 27.9% to 80.1%, and then maintained relative stability. Based on the above results, the optimal pH was determined to be 8.0 for simultaneous removal of volatile phenol and total cyanide in coking wastewater.

The adsorption of AC on phenol and cyanide may not only be related to the pore structure, but also affected by the types of functional groups and the chemical nature of adsorbate. Changes in the pH might alter the electron-withdrawing effect of oxygen-containing functional groups on the surface of SBAC, thereby changing the  $\pi$ - $\pi$  interaction between the pollutant molecules [34]. In addition, under strong alkaline conditions, the enhanced competition of ammonia nitrogen with phenols and cyanide might be another cause of the decrease in removal efficiency.



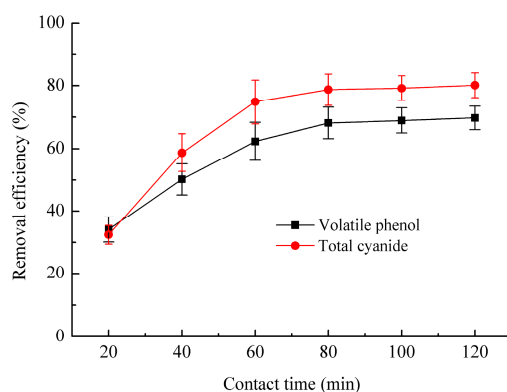
**Figure 4.** Effects of initial pH on the removal efficiencies of volatile phenol and total cyanide. Error bars represent the range of data from duplicate experiments.

#### 3.2.2. Effect of Contact Time

The effect of contact time on the removal efficiencies of volatile phenol and total cyanide in coking wastewater was investigated at a SBAC dosage of 8 g/L, a pH of 8.0, and a temperature of 25 °C. It can be seen from Figure 5 that in the first 60 min, the removal efficiencies of volatile phenol and total



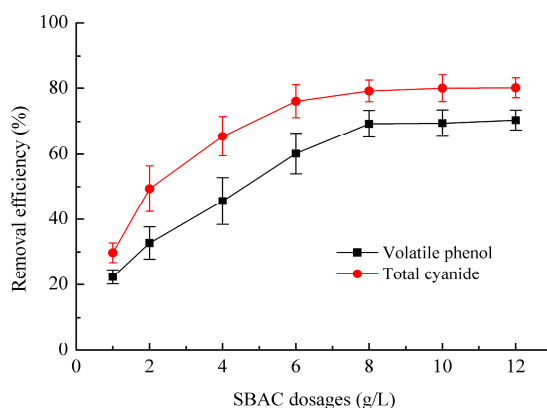
cyanide increased significantly, and remained relatively stable after 80 min. The maximal removal efficiencies of volatile phenol and total cyanide reached 69.7% and 80.1%, respectively. This might be owing to the fact that there were more available active sites on the surface of SBAC in the initial stage of adsorption, and, simultaneously, a higher concentration gradient between wastewater and SBAC was more conducive to the mass transfer of phenol and cyanide to SBAC [12]. However, with the passage of time, the vacant sites and active functional groups on the SBAC reached saturation, leading to the termination of phenol and cyanide adsorption [34].



**Figure 5.** Effects of contact time on the removal efficiencies of volatile phenol and total cyanide. Error bars represent the range of data from duplicate experiments.

### 3.2.3. Effect of SBAC Dosage

The effect of SBAC dosages (1–12 g/L) on the volatile phenol and total cyanide removal in coking wastewater was evaluated at a contact time of 80 min, a pH of 8.0, and a temperature of 25 °C. The result is shown in Figure 6. It can be seen that the removal efficiencies of volatile phenol and total cyanide increased gradually to 69.7% and 80.1%, respectively, with SBAC dosage increasing from 1 to 8 g/L. This phenomenon can be explained as follows: with the increase of SBAC dosage, the provided adsorption surface area and active sites were enhanced [35]. However, no significant changes in volatile phenol and total cyanide were observed when the SBAC dosage exceeded 8 g/L, probably due to the fact that at high SBAC dosage the overlap of some active sites not only reduced the surface area for adsorption, but also decreased the odds of point contacts between SBAC and contaminants [15]. It was also found that the SBAC dosage in this study was higher than other adsorption studies of only phenol or cyanide, which might be attributed to the competitive adsorption of other pollutants in the coking wastewater.



**Figure 6.** Effects of SBAC dosages on the removal efficiencies of volatile phenol and total cyanide. Error bars represent the range of data from duplicate experiments.



### 3.3. Kinetics of Volatile Phenol and Total Cyanide

In this study, pseudo-first-order [36] and pseudo-second-order kinetic models [37] were used for the experimental data to evaluate the adsorption mechanism of volatile phenol or total cyanide onto the SBAC. The linear form of pseudo-first and pseudo-second order kinetic models are shown as Equations (3) and (4), respectively.

$$\ln(q_e - q_t) = \ln q_e - k_1 t \quad (3)$$

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \quad (4)$$

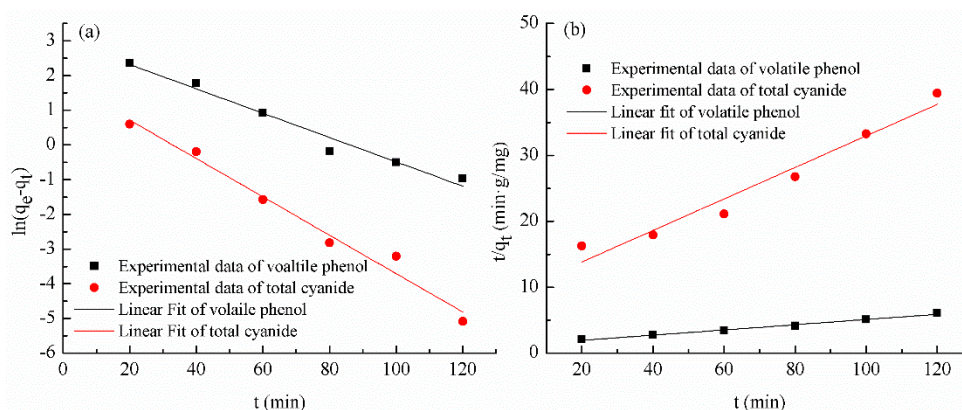
where  $q_t$  is the adsorption amount (mg/g) of volatile phenol or total cyanide at time  $t$  (min), and the  $k_1$  and  $k_2$  are the rate constants of the pseudo-first-order (1/min) and pseudo-second-order adsorption (g/(mg·min)), respectively.

The average relative error (ARE) between the experimental data ( $q_e^{\text{exp}}$ ) and calculated values ( $q_e^{\text{cal}}$ ) was used to validate the goodness of fit of the kinetic models [15]. ARE is calculated according to Equation (5):

$$\text{ARE}(\%) = 100/N \times \sqrt{\sum_{i=1}^P \left( \frac{q_{e,i}^{\text{exp}} - q_{e,i}^{\text{cal}}}{q_{e,i}^{\text{exp}}} \right)^2} \quad (5)$$

where  $N$  is the number of observations in the kinetic experiment, and  $P$  is the number of parameters in the regression model. The smaller ARE values indicate more accurate estimation of  $q_e$  [15].

The linear form of the pseudo-first-order and pseudo-second-order kinetic models for volatile phenol and total cyanide on the SBAC is presented in Figure 7 and Table 4. For the volatile phenol adsorption, both pseudo-first-order and pseudo-second-order kinetic models fitted well with experimental data, and the corresponding correlation coefficients ( $R^2$ ) were 0.974 and 0.991, respectively. However, the calculated  $q_e$  values ( $q_e^{\text{cal}}$ ) obtained from the pseudo-first-order model (20.46 mg/g) were closer to the experimental value (20.02 mg/g) than the value determined by the pseudo-second-order model (25.38 mg/g). This indicated that the pseudo-first-order equation might be more suitable for describing the mechanism of phenol–SBAC interactions. For the total cyanide adsorption, the  $R^2$  of pseudo-first-order was higher than that of pseudo-second-order, but the  $q_e$  value determined from the pseudo-second-order model was more consistent with the experimental value. Therefore, it is proposed that the pseudo-second-order kinetic model should be used to describe the adsorption of cyanides onto SBAC, which is in line with a previous study [34].



**Figure 7.** (a) Pseudo-first-order and (b) pseudo-second-order adsorption kinetics for volatile phenol and total cyanide adsorption onto SBAC.

**Table 4.** Parameters of kinetic modeling of the adsorption of volatile phenol and total cyanide onto SBAC.

Kinetic Modeling Parameters	Pseudo-First-Order				Pseudo-Second-Order				$q_e^{\text{exp}}$
	$k_1$	$q_e^{\text{cal}}$	$R^2$	ARE	$k_2$	$q_e^{\text{cal}}$	$R^2$	ARE	
Volatile phenol	0.0351	20.46	0.9737	1.29	0.0013	25.38	0.9912	25.64	20.02
Total cyanide	0.0553	6.17	0.9804	89.26	0.0063	4.18	0.9605	28.22	3.26

### 3.4. Cost Analysis

In this section, the cost of production per kilogram SBAC is roughly calculated based on raw material fees, reagent fees, water fees, and electricity fees. The one-time investment costs, depreciation costs, freight costs, and labor costs were not included. The industrial water and electricity costs, as well as sludge treatment compensation, were calculated according to the standard in Nanjing, China. The calculation results are described in Table 5. It can be seen that the cost per kilogram of SBAC production was about 0.840 USD, which is cheaper than commercial AC (2.0–2.2 USD/kg) [38]. According to gross cost estimate from previous literature, the cost of sludge-based adsorbents (0.1–0.2 USD/kg) is only 5–10% of commercial AC [39]. In this study, the data used for cost analysis are derived from laboratory studies, and the costs of SBAC production are expected to be reduced through large-scale production.

**Table 5.** Cost analysis of SBAC.

Materials and Energy	Consumption	Unit Cost	Cost (USD)
ZnCl <sub>2</sub> (recovery rate of 70%)	0.315 kg	1.227 (USD/kg)	0.387
HCl	0.169 kg	45.984 (USD/t)	0.008
Water	51.240 L	0.859 (USD/kg)	0.044
Electricity	6.5 kWh	0.104 (USD/kWh)	0.676
BW	0.562 kg	0.359 (USD/kg)	0.202
By-product tar	0.52 kg	308.955 (USD/t)	−0.306
Sludge treatment compensation	7.460 kg	22.992 (USD/t)	−0.171
Total (USD/kg)			0.840

## 4. Conclusions

When SBAC was applied as adsorbent for coking wastewater treatment, volatile phenol and total cyanide were rapidly adsorbed. For the highest removal efficiencies of volatile phenol (69.7%) and total cyanide (80.1%) in the coking wastewater, the optimized SBAC dosage was 8 g/L, the pH was 8.0, and the contact time was 80 min. Due to the addition of BW, a high surface area (289.578 m<sup>2</sup>/g) with the presence of a rich mesoporous structure might be the main reason for the adsorption of volatile phenol and total cyanide over the SBAC. Simultaneously, the increase of oxygen-containing functional groups (-OH, C=O) on the surface of SBAC might be more conducive to driving the chemisorption process of cyanide. In addition, the kinetic model of pseudo-first-order fitted well to the experimental data of volatile phenol, while the adsorption of total cyanide onto SBAC should be described by the pseudo-second-order kinetic model. The cost analysis based on experimental data indicated that the SBAC was more economical than commercial AC, and its production cost might be further reduced by mass production.

**Author Contributions:** Conceptualization: Y.L.; validation: Y.H. and H.C.; investigation: Y.H. and H.C.; data curation: H.C.; writing—original draft preparation: Y.L.; writing—review and editing: Y.L. All authors have read and agreed to the published version of the manuscript.

**Funding:** This work was financially supported by the Nanjing Forestry University Youth Innovation Fund (CX2017025), the National Natural Science Foundation of China (No. 51608272), and the Undergraduate Practical Innovation Training Program Project (2019NFUSPITP0487).

**Conflicts of Interest:** The authors declare no conflict of interest.

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